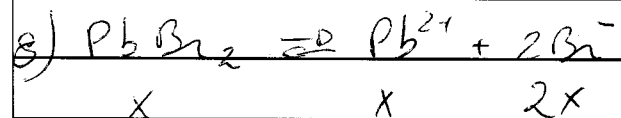


$$K_s = [\text{Mg}^{2+}][\text{OH}^-]^2$$

$$5,61 \times 10^{-12} = x \cdot (2x)^2 = x \cdot 4x^2 = 4x^3$$

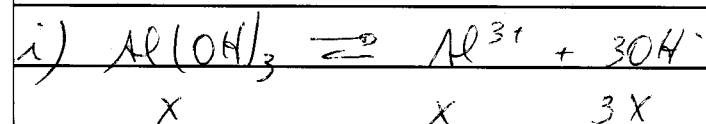
$$x = \sqrt[3]{\frac{5,61 \times 10^{-12}}{4}} = 1,12 \times 10^{-4} \text{ mol l}^{-1}$$



$$K_s = [\text{Pb}^{2+}][\text{Br}^-]^2$$

$$6,60 \times 10^{-6} = x \cdot (2x)^2 = x \cdot 4x^2 = 4x^3$$

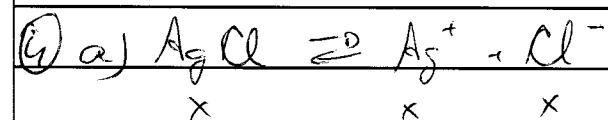
$$x = \sqrt[3]{\frac{6,60 \times 10^{-6}}{4}} = 1,18 \times 10^{-2} \text{ mol l}^{-1}$$



$$K_s = [\text{Al}^{3+}][\text{OH}^-]^3$$

$$3,0 \times 10^{-34} = x \cdot (3x)^3 = x \cdot 9x^3 = 9x^4$$

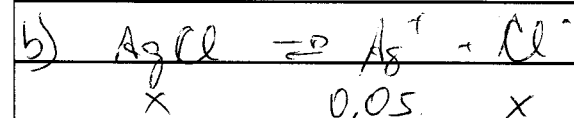
$$x = \sqrt[4]{\frac{3,0 \times 10^{-34}}{9}} = 2,4 \times 10^{-9} \text{ mol l}^{-1}$$



$$K_s = [\text{Ag}^+][\text{Cl}^-]$$

$$1,77 \times 10^{-10} = x \cdot x = x^2$$

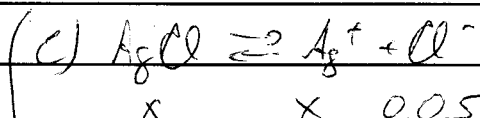
$$x = \sqrt{1,77 \times 10^{-10}} = 1,33 \times 10^{-5} \text{ mol l}^{-1}$$



$$K_s = [\text{Ag}^+][\text{Cl}^-]$$

$$1,77 \times 10^{-10} = 0,05 \cdot x$$

$$x = \frac{1,77 \times 10^{-10}}{0,05} = 3,54 \times 10^{-9} \text{ mol l}^{-1}$$



$$K_s = [\text{Ag}^+][\text{Cl}^-]$$

$$1,77 \times 10^{-10} = x \cdot 0,05$$

$$x = \frac{1,77 \times 10^{-10}}{0,05} = 3,54 \times 10^{-9} \text{ mol l}^{-1}$$