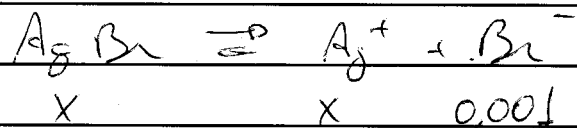


$$K_s = [\text{Ag}^+][\text{Cl}^-]$$

$$1,77 \times 10^{-10} = x \cdot 0,01$$

$$x = \frac{1,77 \times 10^{-10}}{0,01} = 1,77 \times 10^{-8} \text{ mol/L}$$

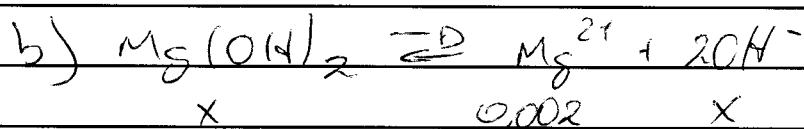
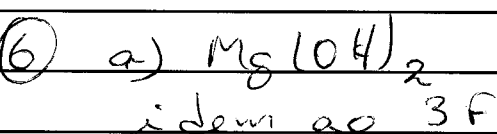


$$K_s = [\text{Ag}^+][\text{Br}^-]$$

$$5,35 \times 10^{-13} = x \cdot 0,001$$

$$x = \frac{5,35 \times 10^{-13}}{0,001} = 5,35 \times 10^{-10} \text{ mol/L}$$

Precipita primeiro o AgBr, pois a $[\text{Ag}^+]$ necessária é menor.



$$K_s = [\text{Mg}^{2+}][\text{OH}^-]^2$$

$$5,61 \times 10^{-12} = x \cdot 0,002^2$$

$$x = \frac{5,61 \times 10^{-12}}{0,0004} = 1,40 \times 10^{-8} \text{ mol/L}$$