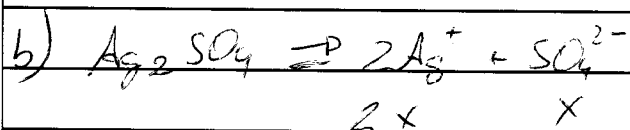


$$K_s = [\text{Ba}^{2+}][\text{SO}_4^{2-}]$$

$$1.0 \times 10^{-10} = x \cdot x = x^2 \Rightarrow x = \sqrt{1.0 \times 10^{-10}} = 1.0 \times 10^{-5} \text{ mol l}^{-1}$$

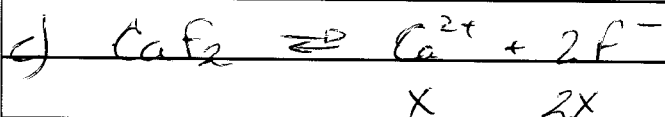


$$2x \quad x$$

$$K_s = [\text{Ag}^+]^2[\text{SO}_4^{2-}]$$

$$1.5 \times 10^{-2} = (2x)^2 \cdot x = 4x^2 \cdot x = 4x^3$$

$$x = \sqrt[3]{\frac{1.5 \times 10^{-2}}{4}} = 0.155 \text{ mol l}^{-1}$$

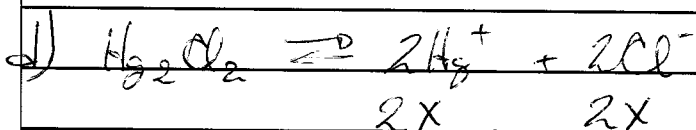


$$x \quad 2x$$

$$K_s = [\text{Ca}^{2+}][\text{F}^-]^2$$

$$4.0 \times 10^{-11} = x \cdot (2x)^2 = x \cdot 4x^2 = 4x^3$$

$$x = \sqrt[3]{\frac{4.0 \times 10^{-11}}{4}} = 2.15 \times 10^{-4} \text{ mol l}^{-1}$$

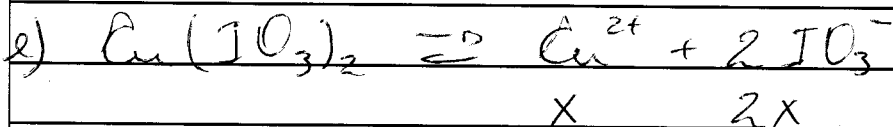


$$2x \quad 2x$$

$$K_s = [\text{Hg}^+]^2[\text{Cl}^-]^2$$

$$1.3 \times 10^{-18} = (2x)^2 (2x)^2 = 4x^2 \cdot 4x^2 = 16x^4$$

$$x = \sqrt[4]{\frac{1.3 \times 10^{-18}}{16}} = 2.39 \times 10^{-5} \text{ mol l}^{-1}$$



$$x \quad 2x$$

$$K_s = [\text{Cu}^{2+}][\text{IO}_3^-]^2$$

$$6.94 \times 10^{-8} = x \cdot (2x)^2 = x \cdot 4x^2 = 4x^3$$

$$x = \sqrt[3]{\frac{6.94 \times 10^{-8}}{4}} = 2.59 \times 10^{-3} \text{ mol l}^{-1}$$