

$$K_c = \frac{[PQ_3] \cdot [Q_2]}{[PQ_5]} \quad K_c = \frac{1,2 \cdot 1,2}{1,8} \quad K_c = 0,8$$

16 $2HI \rightleftharpoons H_2 + I_2$ $K_c = \frac{[H_2] \cdot [I_2]}{[HI]^2}$

5 mol/L	0	0	
-3,0 mol/L	1,5 mol/L	1,5 mol/L	$K_c = \frac{1,5 \cdot 1,5}{(2,0)^2} \quad K_c = 0,561$
2,0	1,5	1,5	

17 $2SO_2 + O_2 \rightleftharpoons 2SO_3$ $K_c = \frac{[SO_3]^2}{[SO_2]^2 \cdot [O_2]}$

6,0 mol/L	5 mol/L	0	
-4,0 mol/L	-2,0 mol/L	4,0 mol/L	$K_c = \frac{(4)^2}{(2)^2 \cdot 3} \quad K_c = 1,33$
2,0	3,0	4,0	

18 $H_2 + I_2 \rightleftharpoons 2HI$ $K_c = \frac{[HI]^2}{[H_2] \cdot [I_2]}$ $K_c = (1,56)^2$

1 mol/L	1 mol/L	0	
-0,78	-0,78	1,56	$K_c = \frac{(1,56)^2}{0,22 \cdot 0,22}$
0,22	0,22	1,56	$K_c = 50,28$

19 $C_4H_8O_2 + H_2O \rightleftharpoons C_2H_6O + C_2H_4O_2$ $V = 2 \text{ litros}$

6,0 mol	6,0 mol	0	0
-4,0	-4,0	4,0	4,0
$\frac{2,0}{2} = 1$	$\frac{2,0}{2} = 1$	$\frac{4,0}{2} = 2$	$\frac{4,0}{2} = 2$ em dois litros

$$K_c = \frac{[C_2H_6O] \cdot [C_2H_4O_2]}{[C_4H_8O_2] \cdot [H_2O]}$$

$$K_c = \frac{2 \cdot 2}{1 \cdot 1} = 4 \quad K_c = 4$$

20 $N_2O_4 \rightleftharpoons 2NO_2$ $K_c = \frac{[NO_2]^2}{[N_2O_4]}$ $K_c = (0,06)^2$

0,75	0	
-0,03	0,06	
0,72	0,06	$K_c = 5,0 \cdot 10^{-3}$