

Determine the type of substance and then name the substance IC = ionic compound
 MC/A = molecular compound/acid E = element MC/BI = molecular compound/binary
 inorganic

IC 1. NaCl *sodium chloride*
 CO₂ *carbon dioxide*

MC/BI 2.

MC/A 3. HNO₃ *nitric acid*
 Cu *copper*

E 4.

E 5. Cl₂ *chlorine*
 Na₃PO₄ *sodium phosphate*

IC 6.

MC/BI 7. H₂SO₃ *sulfurous acid*
 (NH₄)₂SO₄ *ammonium sulfate*

IC 8.

MC/BI 9. N₂O₅ *dinitrogen pentoxide*
 P₄ *phosphorus*

E 10.

MC/A 11. H₃PO₄ *phosphoric acid*
 Ca₃(PO₄)₂ *calcium phosphate*

IC 12.

MC/BI 13. P₄O₁₀ *tetraphosphorus decoxide*
 FeBr₃ *iron (III) bromide*

IC 14.

IC 15. FeBr₂ *iron (II) bromide*
 AgNO₃ *silver nitrate*

IC 16.

MC/BI 17. SF₆ *sulfur hexafluoride*
 S₈ *sulfur*

E 18.

MC/A 19. H₂CO₃ *carbonic acid*
 NaHCO₃ *sodium hydrogen carbonate*

IC 20.

MC/A 21. HCl (aq) *hydrochloric acid*
 (g) *hydrogen chloride*

MC/A 22. HCl

IC 23. Al (OH)₃ *aluminum hydroxide*
 N₂O *dinitrogen monoxide*

MC/BI 24.

MC/A 25. H₂SO₄ *sulfuric acid*
 H₂SO₄ *magnesium nitride*

IC 26.

IC 27. BaI₂ *barium iodide*
 LiOH *lithium hydroxide*

IC 28.

IC 29. NiO *nickel (II) oxide*
N₂O₃ *dinitrogen trioxide*

MC/BI 30.

IC 31. PbS₂ *lead (IV) sulfide*
HC₂H₃O₂ *acetic acid*

MC/A 32.

IC 33. SnSO₃ *tin (II) sulfite*
K₂CO₃ *potassium carbonate*

IC 34.

MC/BI 35. PCl₃ *phosphorus trichloride*
HNO₃ *nitric acid*

MC/A 36.

MC/A 37. HBr (aq) *hydrobromic acid*
(g) *hydrogen bromide*

MC/A 38. HBr

IC 39. Fe₂(SO₄)₃ *iron (III) sulfate*
FeSO₄ *iron (II) sulfate*

IC 40.

IC 41. NaClO₄ *sodium perchlorate*
HClO₄ *perchloric acid*

MC/A 42.