Chemistry
Study Guide Answers: stoichiometry

____N_2 = $3H_2$ ⇢ $2NH_3$

Use ratios of coefficients in the above balanced equation to answer the following:

1. 1 mole N_2 will form __2__ moles NH_3.

2. 8 moles N_2 will form __16__ moles NH_3.

3. 1/2 mole N_2 will form __1__ mole NH_3.

4. .2 mole N_2 will form moles __0.4__ NH_3.

5. 1 mole N_2 will react totally with __3__ moles H_2.

6. 4 moles N_2 will react totally with __12__ moles H_2.

7. 1/3 mole N_2 will react totally with __1__ moles H_2.

8. .2 mole N_2 will react totally with __0.6__ moles H_2.

9. 3 moles of H_2 will form __2__ moles NH_3.

10. 6 moles of H_2 will form __4__ moles NH_3.

11. __1.5__ moles H_2 will form 1 mole NH_3.

12. __12__ moles H_2 will form 8 moles NH_3.

13. 4.2 moles N_2 will form __8.4__ moles NH_3.

14. __46__ moles N_2 will form 92 moles NH_3.

15. 0.4 moles N_2 will totally react with __1.2__ moles H_2

16. 15 moles will totally react with __45__ moles H_2

17. __1/3__ moles N_2 will totally react with 1 mole H_2

18. __8__ moles N_2 will totally react with 24 moles H_2 v

19. __18__ moles H_2 will form 12 moles NH_3

20. __0.3__ moles H_2 will form .2 moles NH_3