

Study Guide: Chemistry, Molar Mass II

$$1.00 \text{ mol CH}_4 = \text{_____ g} = \text{_____ CH}_4$$

$$2.00 \text{ mol CH}_4 = \text{_____ g} = \text{_____ CH}_4$$

$$3.00 \text{ mol CH}_4 = \text{_____ g} = \text{_____ CH}_4$$

$$0.500 \text{ mol CH}_4 = \text{_____ g} = \text{_____ CH}_4$$

$$1.50 \text{ mol CH}_4 = \text{_____ g} = \text{_____ CH}_4$$

$$1.25 \text{ mol CH}_4 = \text{_____ g} = \text{_____ CH}_4$$

$$1.00 \text{ mol O}_2 = \text{_____ g} = \text{_____ O}_2$$

$$0.50 \text{ mol O}_2 = \text{_____ g} = \text{_____ O}_2$$

$$1.50 \text{ mol O}_2 = \text{_____ g} = \text{_____ O}_2$$

$$10.0 \text{ mol O}_2 = 320 \text{ g} = \text{_____ O}_2$$

$$1.00 \text{ mol NaOH} = \text{_____ g} = \text{_____ NaOH}$$

$$3.00 \text{ mol NaOH} = \text{_____ g} = \text{_____ NaOH}$$

$$0.50 \text{ mol NaOH} = \text{_____ g} = \text{_____ NaOH}$$

$$6 \times 10^{23} \text{ N}_2 = \text{_____ mol} = \text{_____ g N}_2$$

$$3 \times 10^{23} \text{ N}_2 = \text{_____ mol N}_2 = \text{_____ g N}_2$$

$$14.0 \text{ g N}_2 = \text{_____ mol N}_2 = \text{_____ N}_2$$

$$280 \text{ g N}_2 = \text{_____ mol N}_2 = \text{_____ N}_2$$

$$1.0 \text{ mol CO}_2 = \text{_____ g} = \text{_____ CO}_2$$

$$88.0 \text{ g CO}_2 = \text{_____ mol CO}_2 = \text{_____ CO}_2$$

$$4.4 \text{ g CO}_2 = \text{_____ mol CO}_2 = \text{_____ CO}_2$$

$$22.0 \text{ g CO}_2 = \text{_____ mol CO}_2 = \text{_____ CO}_2$$

$$1.00 \text{ mol H}_2\text{O}_2 = \text{_____ g H}_2\text{O}_2$$

$$3 \times 10^{23} \text{ H}_2\text{O}_2 = \text{_____ g H}_2\text{O}_2$$

$$68 \text{ g H}_2\text{O}_2 = \text{_____ H}_2\text{O}_2$$

$$340 \text{ g H}_2\text{O}_2 = \text{_____ H}_2\text{O}_2$$

$$12 \times 10^{23} \text{ H}_2\text{O}_2 = \text{_____ g H}_2\text{O}_2$$