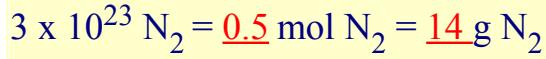
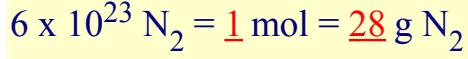
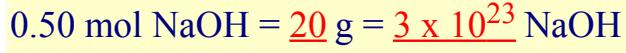
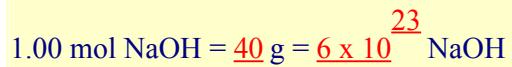
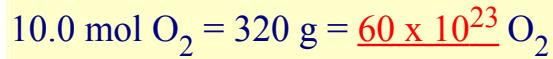
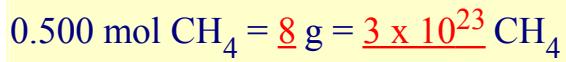
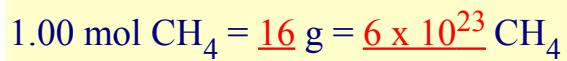


Study Guide: Chemistry, Molar Mass II



$$280 \text{ g N}_2 = \underline{10} \text{ mol N}_2 = \underline{60 \times 10^{23}} \text{ N}_2$$

$$1.0 \text{ mol CO}_2 = \underline{44} \text{ g} = \underline{6 \times 10^{23}} \text{ CO}_2$$

$$88.0 \text{ g CO}_2 = \underline{2} \text{ mol CO}_2 = \underline{12 \times 10^{23}} \text{ CO}_2$$

$$4.4 \text{ g CO}_2 = \underline{0.1} \text{ mol CO}_2 = \underline{0.6 \times 10^{23}} = \underline{6 \times 10^{22}} \text{ CO}_2$$

$$22.0 \text{ g CO}_2 = \underline{0.5} \text{ mol CO}_2 = \underline{3 \times 10^{23}} \text{ CO}_2$$

$$1.00 \text{ mol H}_2\text{O}_2 = \underline{34} \text{ g H}_2\text{O}_2$$

$$\underline{3 \times 10^{23}} \text{ H}_2\text{O}_2 = \underline{17} \text{ g H}_2\text{O}_2$$

$$68 \text{ g H}_2\text{O}_2 = \underline{12 \times 10^{23}} \text{ H}_2\text{O}_2$$

$$340 \text{ g H}_2\text{O}_2 = \underline{60 \times 10^{23}} \text{ H}_2\text{O}_2$$

$$12 \times 10^{23} \text{ H}_2\text{O}_2 = \underline{68} \text{ g H}_2\text{O}_2$$