

Chemistry
Study Guide: stoichiometry



Use ratios of coefficients in the above balanced equation to answer the following:

1. 1 mole N_2 will form ___ moles NH_3 .
2. 8 moles N_2 will form ___ moles NH_3 .
3. $\frac{1}{2}$ mole N_2 will form ___ mole NH_3 .
4. .2 mole N_2 will form ___ moles NH_3 .
5. 1 mole N_2 will react totally with ___ moles H_2 .
6. 4 moles N_2 will react totally with ___ moles H_2 .
7. $\frac{1}{3}$ mole N_2 will react totally with ___ moles H_2 .
8. .2 mole N_2 will react totally with ___ moles H_2 .
9. 3 moles of H_2 will form ___ moles NH_3 .
10. 6 moles of H_2 will form ___ moles NH_3 .
11. ___ moles H_2 will form 1 mole NH_3 .
12. ___ moles H_2 will form 8 moles NH_3 .
13. 4.2 moles N_2 will form ___ moles NH_3 .
14. ___ moles N_2 will form 92 moles NH_3 .
15. .4 moles N_2 will totally react with ___ moles H_2
16. 15 moles will totally react with ___ moles H_2
17. ___ moles N_2 will totally react with 1 mole H_2
18. ___ moles N_2 will totally react with 24 moles H_2
19. ___ moles H_2 will form 12 moles NH_3
20. ___ moles H_2 will form .2 moles NH_3