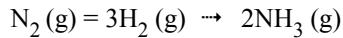


## Chemistry

## Study Guide: Stoichiometry 2



$$1 \text{ mole N}_2 = \underline{\hspace{2cm}} \text{ g N}_2$$

$$1 \text{ mole H}_2 = \underline{\hspace{2cm}} \text{ g H}_2$$

$$1 \text{ mole NH}_3 = \underline{\hspace{2cm}} \text{ g NH}_3$$

$$1. 28 \text{ g N}_2 = \underline{\hspace{2cm}} \text{ moles N}_2 = \underline{\hspace{2cm}} \text{ moles NH}_3 = \underline{\hspace{2cm}} \text{ g NH}_3$$

$$2. 14 \text{ g N}_2 = \underline{\hspace{2cm}} \text{ moles N}_2 = \underline{\hspace{2cm}} \text{ moles NH}_3 = \underline{\hspace{2cm}} \text{ g NH}_3$$

$$3. 280 \text{ g N}_2 = \underline{\hspace{2cm}} \text{ moles N}_2 = \underline{\hspace{2cm}} \text{ moles H}_2 = \underline{\hspace{2cm}} \text{ g H}_2$$

$$4. 140 \text{ g N}_2 = \underline{\hspace{2cm}} \text{ moles N}_2 = \underline{\hspace{2cm}} \text{ moles NH}_3 = \underline{\hspace{2cm}} \text{ g NH}_3$$

$$5. 6 \text{ g H}_2 = \underline{\hspace{2cm}} \text{ moles H}_2 = \underline{\hspace{2cm}} \text{ moles NH}_3 = \underline{\hspace{2cm}} \text{ g NH}_3$$

$$6. 24 \text{ g H}_2 = \underline{\hspace{2cm}} \text{ moles H}_2 = \underline{\hspace{2cm}} \text{ moles NH}_3 = \underline{\hspace{2cm}} \text{ g NH}_3$$

$$7. 2 \text{ g H}_2 = \underline{\hspace{2cm}} \text{ moles H}_2 = \underline{\hspace{2cm}} \text{ moles NH}_3 = \underline{\hspace{2cm}} \text{ g NH}_3$$

$$8. 18 \text{ g H}_2 = \underline{\hspace{2cm}} \text{ moles H}_2 = \underline{\hspace{2cm}} \text{ moles N}_2 = \underline{\hspace{2cm}} \text{ g N}_2$$

$$9. 17 \text{ g NH}_3 = \underline{\hspace{2cm}} \text{ moles NH}_3 = \underline{\hspace{2cm}} \text{ moles N}_2 = \underline{\hspace{2cm}} \text{ g N}_2$$

$$10. 170 \text{ g NH}_3 = \underline{\hspace{2cm}} \text{ moles NH}_3 = \underline{\hspace{2cm}} \text{ moles N}_2 = \underline{\hspace{2cm}} \text{ g N}_2$$

$$1 \text{ mole of a gas @ STP} = \underline{\hspace{2cm}} \text{ L}$$

$$11. 22.4 \text{ L N}_2 = \underline{\hspace{2cm}} \text{ moles N}_2 = \underline{\hspace{2cm}} \text{ moles NH}_3 = \underline{\hspace{2cm}} \text{ L NH}_3$$

$$12. 11.2 \text{ L N}_2 = \underline{\hspace{2cm}} \text{ moles N}_2 = \underline{\hspace{2cm}} \text{ moles NH}_3 = \underline{\hspace{2cm}} \text{ L NH}_3$$

$$13. 134.4 \text{ L H}_2 = \underline{\hspace{2cm}} \text{ moles H}_2 = \underline{\hspace{2cm}} \text{ moles NH}_3 = \underline{\hspace{2cm}} \text{ L NH}_3$$

$$14. 33.6 \text{ L N}_2 = \underline{\hspace{2cm}} \text{ moles N}_2 = \underline{\hspace{2cm}} \text{ moles H}_2 = \underline{\hspace{2cm}} \text{ L H}_2$$

$$15. 2.24 \text{ L NH}_3 = \underline{\hspace{2cm}} \text{ moles NH}_3 = \underline{\hspace{2cm}} \text{ moles N}_2 = \underline{\hspace{2cm}} \text{ L N}_2$$

$$1 \text{ mole} = \underline{\hspace{2cm}} \text{ particles}$$

$$16. 6 \times 10^{23} \text{ molecules N}_2 = \underline{\hspace{2cm}} \text{ moles N}_2 = \underline{\hspace{2cm}} \text{ moles H}_2 = \underline{\hspace{2cm}} \text{ molecules H}_2$$

$$17. 3 \times 10^{23} \text{ molecules N}_2 = \underline{\hspace{2cm}} \text{ moles N}_2 = \underline{\hspace{2cm}} \text{ moles NH}_3 = \underline{\hspace{2cm}} \text{ molecules NH}_3$$

$$18. 36 \times 10^{23} \text{ molecules H}_2 = \underline{\hspace{2cm}} \text{ moles H}_2 = \underline{\hspace{2cm}} \text{ moles NH}_3 = \underline{\hspace{2cm}} \text{ molecules NH}_3$$

$$19. 60 \times 10^{23} \text{ molecules N}_2 = \underline{\hspace{2cm}} \text{ moles N}_2 = \underline{\hspace{2cm}} \text{ moles H}_2 = \underline{\hspace{2cm}} \text{ molecules H}_2$$

$$20. 15 \times 10^{23} \text{ molecules N}_2 = \underline{\hspace{2cm}} \text{ moles N}_2 = \underline{\hspace{2cm}} \text{ moles H}_2 = \underline{\hspace{2cm}} \text{ molecules H}_2$$

$$21. 56 \text{ g N}_2 = \underline{\hspace{2cm}} \text{ g H}_2$$

$$22. 44.8 \text{ L N}_2 = \underline{\hspace{2cm}} \text{ L NH}_3$$

$$23. 12 \times 10^{23} \text{ molecules N}_2 = \underline{\hspace{2cm}} \text{ molecules H}_2$$

2 mole

6 mole